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ICSE/Chemistry/Std 9/ch 3- Solubility Curve \u0026amp; Application|| Explanation of solubility curve|| Water Unit 12 Solutions Solubility Curves

UNIT 12 - SOLUTIONS Chemistry-Wilson 2 Solubility Curves:

1. Which compound is least soluble at: (A) 20°C (B) 80°C 2. Which substance is most soluble at: (A) 10°C (B) 50°C (C) 90°C 3. The solubility of which substance is most affected by changes in temperature? 4. The solubility of which substance is least affected by changes in ...

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UNIT 12 REVIEW WORKSHEET Part 1 □ Solubility Curves - USE YOUR SOLUBILITY CURVE GRAPH TO ANSWER

#1-4. 1. At what temp does 135 grams of KI dissolved in 100

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grams of water form a saturated solution? 2. o How many grams of  $\text{KNO}_3$  will dissolve in 400 grams of water at  $60^\circ\text{C}$ ? 3. If 10 grams of  $\text{KClO}_3$  are dissolved in 100 grams of water at  $30^\circ\text{C}$

## SOLUBILITY CURVES WORKSHEET

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Which compound is least soluble at  $20^\circ\text{C}$ ? At  $80^\circ\text{C}$

Solubility Worksheet 2 - UNIT 12 SOLUTIONS SOLUBILITY

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UNIT 12 REVIEW WORKSHEET Part 1 □ Solubility Curves -

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USE YOUR SOLUBILITY CURVE GRAPH TO ANSWER

#1-4. 1. At what temperature does 135 grams of KI dissolved in 100 grams of water form a saturated solution? 2. How many grams of  $\text{KNO}_3$  will dissolve in 400 grams of water at  $60^\circ\text{C}$ ? 3. If 10 grams of  $\text{KClO}_3$  are dissolved in 100 grams of water at ...

## SOLUBILITY CURVES WORKSHEET

Use your solubility curve graphs provided to answer the following questions. 1. What are the customary units of solubility on solubility curves? Degrees Celsius and grams of solute/100g of water 2. Define solubility. A measure of how much solute can dissolve in a given amount of solvent. 3. According to the graph, the solubility of any ...

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Answers of water 2. Define solubility. A measure of how much solute can dissolve in a given amount of solvent. 3. According to the graph, the solubility of any substance changes as

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Interpreting Solubility Curves 3. What will happen to this solute when 12 g is added to 100 g of water at 200C? 4. What type of solution is obtained when 12 g of this solute is added to 100 g of water at 200C (unsaturated, saturated, or supersaturated)? 5. At 200C, what is the maximum amount of this solute that can be dissolved in 100 g of ...

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Supersaturated: A solution that has an excess of solute beyond the solubility point for a given temperature. Excess solute will either precipitate out, or remain in an unstable dissolved state until the supersaturated solution is disturbed.

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Unit 10: Solutions-Key Regents Chemistry 14 Mr. Murdoch

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Solubility (g solute/100. g H<sub>2</sub>O) 150. 140. 130. 120. 110.  
100. 90. 80. 70. 60. 50. 40. 30. 20. 10. 0 KI NaNO<sub>3</sub> KNO<sub>3</sub>  
NH<sub>4</sub>Cl HCl KCl NaCl KClO<sub>3</sub> NH<sub>3</sub> SO<sub>2</sub> Table G Solubility  
Curves at Standard Pressure

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